


Name:													
Enrolment No:													
<b>UPES</b> <b>End Semester Examination, May 2023</b>													
<b>Course: Process Modelling and Simulation</b> <b>Program: M. Tech Chemical Engineering</b> <b>Course Code: CHPD7009</b>		<b>Semester : II</b> <b>Time : 03 hrs.</b> <b>Max. Marks: 100</b>											
<b>Instructions: 1) Answer the questions section wise in the answer booklet. 2) Assume suitable data wherever necessary. 3) The notations used here have the usual meanings.</b>													
<b>SECTION A</b> <b>(5Qx4M=20Marks)</b>													
S. No.		Marks	CO										
Q 1	Discuss about dynamic modelling of a chemical engineering system.	04	CO1										
Q 2	Explain about the mathematical consistency of a process model.	04	CO1										
Q 3	Discuss the sequential modular approach of process simulation	04	CO1										
Q 4	State the assumptions of continuously stirred tank reactor.	04	CO1										
Q 5	Define the variables and the parameters used in a mathematical model.	04	CO1										
<b>SECTION B</b> <b>(4Qx10M= 40 Marks)</b>													
Q 6	Illustrate the application of Neumann and Dirichlet boundary conditions in mathematical modelling with suitable examples.	10	CO2										
Q 7	For laminar flow, the friction coefficient ' $f$ ' is related to Reynolds number as $f = aRe^b$ . Determine $a$ & $b$ using a method of least squares to fit the following data.	10	CO2										
	<table border="1" style="margin-left: auto; margin-right: auto;"> <tr> <td><math>f</math></td> <td>500</td> <td>1000</td> <td>1500</td> <td>2000</td> </tr> <tr> <td><math>Re</math></td> <td>0.0320</td> <td>0.0160</td> <td>0.0107</td> <td>0.0080</td> </tr> </table>	$f$	500	1000	1500	2000	$Re$	0.0320	0.0160	0.0107	0.0080		
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$Re$	0.0320	0.0160	0.0107	0.0080									
Q 8	Write down the mass balance and energy balance equations for multicomponent distillation column for single stage operation.	10	CO3										
Q 9	Describe the solution of partial differential equations by Orthogonal collocation technique.  <b><u>OR</u></b>  Classify the partial differential equations.	10	CO5										

**SECTION-C**  
**(2Qx20M=40 Marks)**

Q 10	Explain the application of a Newton-Raphson's method to solve the mathematical model equations of a triple-effect evaporator system for a forward feed arrangement.	<b>20</b>	<b>CO3</b>
Q 11	<p>Develop a two-dimensional pseudo heterogeneous model of fixed bed reactor with suitable boundary conditions. State the assumptions clearly.</p> <p style="text-align: center;"><b><u>OR</u></b></p> <p>The Van de Vusse reaction operated in a continuous stirred tank reactor is given as follows:</p> $A \xrightarrow{k_1} B \xrightarrow{k_2} C$ $2A \xrightarrow{k_3} D$ <p>The component material balance equations are described as:</p> $\frac{dC_A}{dt} = -k_1 C_A - k_3 C_A^2 + \frac{F}{V} (C_{Af} - C_A)$ $\frac{dC_B}{dt} = k_1 C_A - k_2 C_B + \frac{F}{V} C_B$ <p>(i) Compute the steady state values of <math>C_A</math> and <math>C_B</math> using the given data.</p> <p>(ii) Perform three iterations employing the fourth-order Runge–Kutta method for dynamic study using a step size of 1 for <math>0 \leq t \leq 2</math> h for a given data and plot the values of <math>C_A</math> and <math>C_B</math> as a function of time.</p> <p>Data:  Reactor volume (<math>V</math>) = 2 lit  Feed flow rate (<math>F</math>) = 50 lit/h  Feed concentration of reactant <math>A(C_{Af})</math> = 10 mol/lit  Kinetic constant (<math>k_1</math>) = 50 h<sup>-1</sup>  Kinetic constant (<math>k_2</math>) = 100 h<sup>-1</sup>  Kinetic constant (<math>k_3</math>) = 10 lit/mol.h</p>	<b>20</b>	<b>CO4</b>